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Wei Kong Pang

*Curtin University of Technology*, [wkpang@uow.edu.au](mailto:wkpang@uow.edu.au)

It-Meng Low

*Curtin University of Technology*

B H. O'Connor

*Curtin University of Technology*

Andrew J. Studer

*Australian Nuclear Science And Technology Organisation*

Vanessa Peterson

*ANSTO*, [vep@ansto.gov.au](mailto:vep@ansto.gov.au)

*See next page for additional authors*

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## Comparison of thermal stability in MAX211 and 312 phases

### Abstract

The susceptibility of four MAX phases (Ti<sub>2</sub>AlC, Cr<sub>2</sub>AlC, Ti<sub>3</sub>AlC<sub>2</sub>, and Ti<sub>3</sub>SiC<sub>2</sub>) to high-temperature thermal dissociation in vacuum has been investigated using in-situ neutron diffraction.

### Keywords

312, thermal, comparison, stability, phases, max211

### Disciplines

Engineering | Science and Technology Studies

### Publication Details

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### Authors

Wei Kong Pang, It-Meng Low, B H. O'Connor, Andrew J. Studer, Vanessa Peterson, Zheng-Ming Sun, and J P. Palmquist

## Comparison of thermal stability in MAX 211 and 312 phases

W K Pang<sup>1</sup>, I M Low<sup>1</sup>, B H O'Connor<sup>1</sup>, A J Studer<sup>2</sup>, V K Peterson<sup>2</sup>, Z M Sun<sup>3</sup>  
and J -P Palmquist<sup>4</sup>

<sup>1</sup> Centre for Materials Research, Department of Imaging and Applied Physics, Curtin University of Technology, GPO Box U 1987, Perth WA, Australia.

<sup>2</sup> The Bragg Institute, ANSTO, PMB 1, Menai, NSW 2234, Australia

<sup>3</sup> Materials Research Institute for Sustainable Development, National Institute of Advanced Industrial Science and Technology (AIST), Nagoya 463-8560, Japan.

<sup>4</sup> Kanthal AB, P.O. Box 502, SE-734 27 Hallstahammar, Sweden.

E-mail: j.low@curtin.edu.au

**Abstract.** The susceptibility of four MAX phases ( $\text{Ti}_2\text{AlC}$ ,  $\text{Cr}_2\text{AlC}$ ,  $\text{Ti}_3\text{AlC}_2$ , and  $\text{Ti}_3\text{SiC}_2$ ) to high-temperature thermal dissociation in vacuum has been investigated using *in-situ* neutron diffraction. In high vacuum, these phases decomposed above 1400°C through the sublimation of *M* and *A* elements, forming a surface coating of MC. The apparent activation energies for the decomposition of sintered  $\text{Ti}_3\text{SiC}_2$ ,  $\text{Ti}_3\text{AlC}_2$ , and  $\text{Ti}_2\text{AlC}$  were determined to be 179.3, -71.9, and 85.7 kJ mol<sup>-1</sup>, respectively. The spontaneous release of  $\text{Ti}_2\text{AlC}$  and TiC from de-intercalation during decomposition of  $\text{Ti}_3\text{AlC}_2$  resulted in a negative activation energy.

### 1. Introduction

A group of layered ternary ceramics with the general formula  $M_{n+1}AX_n$  ( $n = 1-3$ , *M* is an early transition metal, *A* is a group A element, and *X* is either carbon and/or nitrogen) were found to be excellent candidate materials for high temperature applications, due to the unique combination of both metallic and ceramic properties [1-6]. MAX phases are layered hexagonal (space group  $P6_3/mmc$ )  $Z = 2$ . In brief, the  $M_6X$  octahedra are edge sharing and are identical to the *NaCl*-type structure of the corresponding binary carbides. These  $M_6X$  octahedra are interleaved by close-packed planes of the *A* element [7-10]. Ternary carbides such as  $\text{Ti}_3\text{AlC}_2$ ,  $\text{Ti}_3\text{SiC}_2$ ,  $\text{Ti}_2\text{AlC}$ , and  $\text{Cr}_2\text{AlC}$  are common representatives of the MAX phases. The electrical and mechanical properties, and the oxidation of these four compounds has been well studied [9, 11-21], however, very little is published on their susceptibility to high-temperature phase dissociation in vacuum.

In this paper, the thermal stability and the corresponding phase dissociation of these ternary carbides in vacuum was investigated using high-temperature *in-situ* neutron diffraction. The apparent activation energies of phase decomposition were calculated. The characteristics of thermal stability in MAX 211 ( $\text{Ti}_2\text{AlC}$  and  $\text{Cr}_2\text{AlC}$ ) and 312 phases ( $\text{Ti}_3\text{AlC}_2$  and  $\text{Ti}_3\text{SiC}_2$ ) are compared and discussed.

### 2. Experimental

The collection of high-temperature *in-situ* neutron diffraction data were collected using Wombat (the high-intensity neutron powder diffractometer) at the OPAL source in Australia.. Data were collected using neutrons with incident wavelength of  $\lambda = \sim 1.660 \text{ \AA}$  from 15 to 135° 2 $\theta$ , with a step size of 0.125°, with the use of the oscillating tertiary collimator. Rietica 1.7.7 was used for phase

identification and Rietveld refinement. The optimized parameters during refinement were background coefficients, zero-shift error, peak shape parameters, cell parameters, and anisotropic thermal factors. The residual values of the refinement, statistical reliability factor of Bragg ( $R_B$ ), R-weighted pattern ( $R_{wp}$ ), R-expected ( $R_{exp}$ ), and the goodness-of-fit ( $\chi^2$ ), were evaluated. In Rietica,  $\chi^2$  is defined as the square of the ratio of  $R_{wp}$  to  $R_{exp}$ .

Fully-dense solid bars of hot-isostatically-pressed  $Ti_2AlC$ ,  $Cr_2AlC$ ,  $Ti_3SiC_2$ , and spark-plasma sintered  $Ti_3AlC_2$  were used in this study [2, 16]. The temperature of sample environment was controlled by a closed cylindrical niobium vacuum furnace ( $10^{-6}$ - $10^{-8}$  torr). The sample was held by vanadium wire and heated to  $1000^\circ C$  at a heating rate of  $10^\circ C/min$  and thereafter at  $5^\circ C/min$  to  $1550^\circ C$ . The dwell times at between  $1000$  and  $1550^\circ C$  are shown in Table 1. Diffraction patterns were collected every minute from starting to the end of the experiment.

**Table 1.** Heating schedule during the in-situ neutron diffraction study.

Temperature ( $^\circ C$ )	Dwell time (min)	Temperature ( $^\circ C$ )	Dwell time (min)
1000	30	1400	200
1100	30	1450	200
1200	30	1500	200
1300	30	1550	200

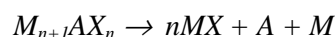
The decomposition rates at different temperatures were calculated and the corresponding apparent activation energies were determined in order to understand further the mechanism of high-temperature phase dissociation in vacuum.

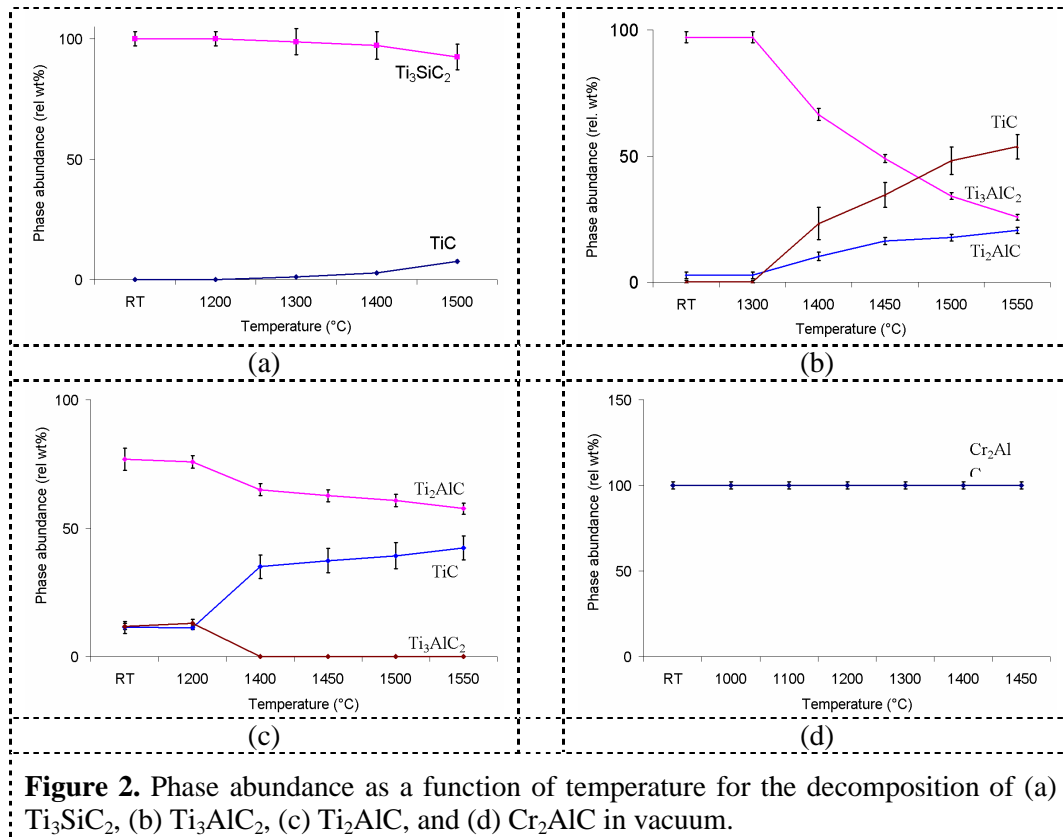
### 3. Results and Discussion

#### 3.1 Phase Transitions during Thermal Dissociation of MAX 211 and 312 Phases

The phase evolution of MAX phases and the relative phase abundances at various temperatures as revealed by in-situ neutron diffraction is shown in Figure 2. The samples used in this experiment were not single-phase with TiC as the most common impurity. The starting compositions of each sample are shown in Figure 2. For  $Ti_3SiC_2$ , it decomposed to TiC above  $1300^\circ C$  (Fig. 2a). A weight loss of  $\sim 4\%$  was observed for decomposed  $Ti_3SiC_2$  which may be attributed to the release of gaseous Ti and Si by sublimation during the decomposition process. These results agree with those reported by Low and co-workers [22-23]. For  $Ti_3AlC_2$ , its decomposition into TiC and  $Ti_2AlC$  as lower order or intermediate phase was observed at  $\geq 1400^\circ C$  (Fig. 2b). However, at higher temperatures, when compared to TiC, a smaller growth rate for  $Ti_2AlC$  may indicate that  $Ti_2AlC$  experienced further decomposition into TiC via the sublimation of Al and Ti, similar to decomposition of  $Ti_3SiC_2$ . In contrast to  $Ti_3AlC$ , no intermediate or lower order phase was observed for the decomposition of  $Ti_3SiC_2$ . This difference can be attributed to the fact that  $Ti_3SiC_2$  is the only stable ternary phase in Ti-Si-C system. For  $Ti_2AlC$  (Fig. 2c), the formation of TiC results mainly from the decomposition of  $Ti_3AlC_2$  as the impurity phase.  $Ti_2AlC$  would only decompose above  $1450^\circ C$  by sublimation of Ti and Al. Figure 2(d) shows the excellent stability of  $Cr_2AlC$  up to  $1450^\circ C$  where no observable phase dissociation was detected. However, it started to react with the vanadium wire and became unstable above  $1450^\circ C$ . Like  $Ti_2AlC$ , it would only decompose above  $1450^\circ C$  via the sublimation of Cr and Al.

The existence of a lower order 211( $Ti_2AlC$ ) phase in  $Ti_3AlC_2$  is believed to be responsible for its unique two-step decomposition behaviour when compared to  $Ti_2AlC$ ,  $Cr_2AlC$ ,  $Ti_3SiC_2$  which do not possess such a lower order phase. As a consequence, the former undergoes an initial spontaneous decomposition to lower order  $Ti_2AlC$  phase, followed by eventual decomposition of  $Ti_2AlC$  through the sublimation of Al and Ti at higher temperature. On the contrary, the latter decompose at elevated temperature in vacuum through the direct sublimation of M and A elements with the concomitant release of a binary carbide (MX) as follows:





**Figure 2.** Phase abundance as a function of temperature for the decomposition of (a)  $\text{Ti}_3\text{SiC}_2$ , (b)  $\text{Ti}_3\text{AlC}_2$ , (c)  $\text{Ti}_2\text{AlC}$ , and (d)  $\text{Cr}_2\text{AlC}$  in vacuum.

### 3.2 Apparent Activation Energies During Isothermal Decomposition

During the isothermal decomposition of 211 and 312 phases at different temperatures, the decomposition rates were calculated by analyzing the relationship between the phase abundances and soaking time. These rates were used to determine the corresponding apparent activation energy of decomposition using an Arrhenius equation [24],  $k = A \exp(-E / RT)$ , where  $k$  is the reaction rate,  $E$  is activation energy,  $R$  is gas constant,  $T$  is absolute temperature and  $A$  is a constant.

The calculated apparent activation energies for the four MAX phases and the proposed reactions are summarized and listed in Table 1.

**Table 1.** Comparison of the decomposition of MAX 211 and 312 phases.

MAX phase	Apparent activation energy	Proposed reactions
$\text{Ti}_3\text{SiC}_2$	$179.3 \text{ kJ mol}^{-1}$	$\text{Ti}_3\text{SiC}_{2(s)} \rightarrow 2\text{TiC}_{(s)} + \text{Ti}_{(g)} + \text{Si}_{(g)}$ (1)
$\text{Ti}_3\text{AlC}_2$	$-71.9 \text{ kJ mol}^{-1}$	$\text{Ti}_3\text{AlC}_{2(s)} \rightarrow \text{Ti}_2\text{AlC}_{(s)} + \text{TiC}_{(s)}$ (2a)
		$\text{Ti}_2\text{AlC}_{(s)} \rightarrow \text{TiC}_{(s)} + \text{Ti}_{(g)} + \text{Al}_{(g)}$ (2b)
$\text{Ti}_2\text{AlC}$	$85.7 \text{ kJ mol}^{-1}$	$\text{Ti}_2\text{AlC}_{(s)} \rightarrow \text{TiC}_{(s)} + \text{Ti}_{(g)} + \text{Al}_{(g)}$ (3)
$\text{Cr}_2\text{AlC}$	Stable in vacuum up to 1450 °C.	$2\text{Cr}_2\text{AlC}_{(s)} \rightarrow \text{Cr}_3\text{C}_{2(s)} + \text{Cr}_{(g)} + 2\text{Al}_{(g)}$ (4)

A negative activation energy obtained for  $\text{Ti}_3\text{AlC}_2$  can be attributed to the presence of weak Ti-Al bonds [25] which will favour spontaneous decomposition through the de-intercalation of TiC from the crystal structure (see Eq 2a). Subsequent decomposition of released  $\text{Ti}_2\text{AlC}$  at higher temperature (see

Equation 2b) would involve a small but positive activation energy. It follows that the high activation energy obtained for  $\text{Ti}_3\text{SiC}_2$  can be attributed to much stronger bonding between Ti and Si, leading to better resistance against thermal dissociation *via* sublimation of Ti and Si from the crystal structure. The lower but positive activation energy observed for decomposition of  $\text{Ti}_2\text{AlC}$  (see Eq. 3) is related to weaker Ti-Al bonds and thus its susceptibility to sublimation in vacuum at elevated temperature.

#### 4. Conclusions

The high-temperature thermal stability of MAX 211 and 312 phases in vacuum has been studied using in-situ neutron diffraction. Both 211 and 312 phases were susceptible to decomposition above 1400°C through sublimation of M and A elements, resulting in a surface coating of TiC being formed. The apparent activation energies for the decomposition of sintered  $\text{Ti}_3\text{SiC}_2$ ,  $\text{Ti}_3\text{AlC}_2$  and  $\text{Ti}_2\text{AlC}$  were determined to be 179.3, -71.9 and 85.7 kJ mol<sup>-1</sup>, respectively. The negative activated energy for  $\text{Ti}_3\text{AlC}_2$  can be attributed to the release of lower order  $\text{Ti}_2\text{AlC}$  and TiC from the crystal structure via spontaneous de-intercalation.

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#### References

- [1] Barsoum M W 2000 *Prog. Solid State Chem.* **28** 201.
- [2] Barsoum M W, Brodtkin D and El-Raghy T 1997 *Scripta Mater.* **36** 535.
- [3] Barsoum M W, Buschow K H J, Robert W C, Merton C F, Bernard I, Edward J K, Subhash M and Patrick V in "Encyclopedia of Materials: Science and Technology" (Elsevier, Oxford, 2006) p. 1.
- [4] Barsoum M W, Crossley A and Myhra S 2002 *J. Phys. Chem. Solids* **63** 2063.
- [5] Barsoum M W and El-Raghy T 2001 *Am. Sci.* **89** 334.
- [6] Barsoum M W, Buschow K H J, Robert W C, Merton C F, Bernard I, Edward J K, Subhash M and Patrick V in "Encyclopedia of Materials: Science and Technology" (Elsevier, Oxford, 2004) p. 1.
- [7] Wu L, Chen J X, Liu M Y, Bao Y W and Zhou Y C 2009 *Wear* **266** 158.
- [8] Wang J, Zhou Y, Lin Z and Hu C 2008 *Scripta Mater.* **58** 1043.
- [9] Wang J, Zhou Y, Liao T, Zhang J and Lin Z 2008 *Scripta Mater.* **58** 227.
- [10] Zhang H Z and Wang S Q 2007 *Acta Mater.* **55** 4645.
- [11] Han J H, Hwang S S, Lee D and Park S W 2008 *J. Eur. Ceram. Soc.* **28** 979.
- [12] Yang C, Jin S, Liang B, Liu G, Duan L and Jia S 2009 *J. Alloys Compd.* **472** 79.
- [13] Yang C, Jin S, Liang B, Liu G and Jia S 2009 *J. Mater. Process. Tech.* **209** 871.
- [14] Lee D B, Nguyen T D, Han J H and Park S W 2007 *Corr. Sci.* **49** 3926.
- [15] Lin Z J, Li M S, Wang J Y and Zhou Y C 2007 *Acta Mater.* **55** 6182.
- [16] Tian W, Sun Z, Du Y and Hashimoto H 2008 *Mater. Lett.* **62** 3852.
- [17] Fraczkiwicz M, Zhou A G and Barsoum M W 2006 *Acta Mater.* **54** 5261.
- [18] Kisi E H, Crossley J A A, Myhra S and Barsoum M W 1998 *J. Phys. Chem. Solids* **59** 1437.
- [19] Kooi B J, Poppen R J, Carvalho N J M, De Hosson J T M and Barsoum M W 2003 *Acta Mater.* **51** 2859.
- [20] Khoptiar Y and Gotman I 2002 *Mater. Lett.* **57** 72.
- [21] Wang P, Mei B C, Hong X L and Zhou W B 2007 *Trans. Nonferrous Metals Soc. Chi.* **17** 1001.
- [22] Low I M 2004 *Mater. Lett.* **58** 927.
- [23] Low I M, Oo Z and Prince K E 2007 *J. Am. Ceram. Soc.* **90** 2610.
- [24] Riley D P, Kisi E H and Hansen T C 2008 *J. Am. Ceram. Soc.* **91** 3207.
- [25] Xu X, Wu E, Du X, Tian Y, and He J 2008 *J. Phys. Chem. Solids* **69** 1356.