Substituted ferrocenes and iodine as synergistic thermoelectrochemical heat harvesting redox couples in ionic liquids

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Abstract
2016 The Royal Society of Chemistry. Combining ferrocene and iodine results in enhanced thermoelectrochemical (or thermogalvanic) waste heat harvesting abilities, for both the Seebeck coefficient and the overall power output. All systems displayed a mixture of ferrocene, ferrocenium, iodine and triiodide. The observed enhancement correlates with lower electron-density on the ferrocene; the synergistic improvement observed for mixtures of substituted ferrocenes and iodine is attributed to the formation of charge-transfer complexes. Combining dibutanoylferrocene and iodine resulted in the highest Seebeck coefficient of 1.67 mV K⁻¹.

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Substituted ferrocenes and iodine as synergistic thermoelectrochemical heat harvesting redox couples in ionic liquids†

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Combining ferrocene and iodine results in enhanced thermoelectrochemical (or therrogalvanic) waste heat harvesting abilities, for both the Seebeck coefficient and the overall power output. All systems displayed a mixture of ferrocene, ferrocinium, iodine and triiodide. The observed enhancement correlates with lower electron-density on the ferrocene; the synergistic improvement observed for mixtures of substituted ferrocenes and iodine is attributed to the formation of charge-transfer complexes. Combining dibutanylfereconcene and iodine resulted in the highest Seebeck coefficient of 1.67 mV K−1.

Worldwide, heat is converted into electricity on a tremendous scale. The Seebeck effect allows arguably the simplest process, whereby a difference in temperature can be converted directly into electrical power with no moving parts.1 This simplicity has made it a hallmark energy system of unmanned space exploration.1 The majority of thermoelectric systems are based upon solid-state semiconductor devices, typically containing rare metalloids such as Bi2Te3,2 and which suffer from rigidity, expense and low efficiencies.1

Thermoelectrochemical cells containing a liquid electrolyte with a suitable redox couple can also generate electrical power from a temperature difference. These systems offer many advantages such as flexibility,3 low cost4 and are suited to the harvesting of low-grade heat.5

The Seebeck coefficient (See, eqn (1)) expresses the potential difference (∆E) generated across two electrodes sharing a common electrolyte, as a function of their temperature difference (∆T).6 Dissolved redox couples at different temperatures will have a free-energy difference proportional to the entropy difference between the redox states (∆S). Unlike in semi-conductors, these liquid, solid or gaseous electrochemical systems give rise to significant See values on the order of mV K−1.7 This is important, as the maximum power is directly proportional to the See value, and the efficiency of heat conversion is proportional to the See squared.8

\[
S_e = \frac{\Delta E}{\Delta T} = \frac{\Delta S}{nF}
\]  

(1)

The majority of studies into thermoelectrochemical cells have employed aqueous electrolytes containing inorganic redox couples. A prime example is aqueous potassium ferrocyanide/ferricyanide solutions, which exhibit a relatively large See on the order of ca. −1.4 mV K−1.3,4 Short-lived systems can exceed this, recently reaching +13.6 mV K−1.9 However, the volatility of water can be a major limitation for thermoelectrochemical cells and so ionic liquids (ILs) have been proposed as non-volatile electrolytes for thermoelectrochemical cells.5 IL-based thermoelectrochemical cells can be operated at temperatures above 100 °C with no degradation, as a result of the high thermal stability and the low (near-zero) vapour pressure of many ILs.5a ILs are known to be highly structured liquids, and entropy is highly significant in many processes in ILs.10 In 2013 a Co(II)/Co(III) complex reached See values of 1.40 to 1.88 mV K−1 in a range of ILs.5b

Despite progress, relying upon the entropy change inherent in a single redox process is clearly self-limiting; too great an entropy change generally correlates with lower exchange current densities and necessitates an electrocatalytic material. Electrocatalysis in ILs is a relatively new and poorly understood area.11 In an effort to move beyond this limiting factor, this work investigates combining two redox couples based upon ferrocene and iodine. When employed individually they represent poor thermoelectrochemical redox couples, but together they display a synergistically enhanced performance. This opens new routes to maximising performance by overcoming the limitations noted above.

Both the ferrocene/ferrocinium12 and iodide/triiodide13 redox couples find extensive uses in electrochemistry, and have been (separately) well characterised in ILs. Both ferrocenium-based ILs14 and iodide/triiodide-based ILs13 are known. Recently, combining
redox couples has been beneficially employed in IL-based dye sensitised solar cells.\textsuperscript{15} In this work, redox couples were investigated by dissolving individual redox couples (ferrocene|ferroenium and iodide|triiodide) and also by mixing ferrocene and iodide, which are in equilibrium with their ferrocenium triiodide salt (eqn (2)). The reaction between ferrocene and iodide has been the subject of extensive fundamental research relating to equilibrium constants and kinetics,\textsuperscript{16} although not in ILs.

In addition to the ferrocenium triiodides being inherently electroactive, the overall equilibrium is also temperature-sensitive. While ferrocene species rapidly decomposes in oxygenated organic solvents,\textsuperscript{17} they were found to be stable when dissolved in the ionic liquid 1-ethyl-3-methyl-imidazolium bis(trifluoromethylsulfonyl)imide ([Emim][NTf\textsubscript{2}]), hence all studies were performed in this IL.

Several substituted ferrocenes were synthesised (full details in ESI). They are ferrocene (Fc), acetylferrocene (AcFc), ethylferrocene (EtFc), dibutylferrocene (DiBuFc) and dibutanoylferrocene (DiBoylFc).

The Seebeck coefficients \(S_e\) were measured in a U-shaped assembly (U-cell, full details in the ESI) using platinum electrodes. Fig. 1(a) displays the potential difference of 15 mM Fc/15 mM Fc\textsuperscript{+} (○), 15 mM Fc/15 mM [I\textsubscript{3}]\textsuperscript{-} (△) and 30 mM Fc/30 mM Fc\textsuperscript{+} ( ■) in [Emim][NTf\textsubscript{2}], measured as a function of temperature; the corresponding \(S_e\) values are summarised in Table 1.

The value measured here for Fc/Fc\textsuperscript{+} is essentially identical to the value reported by Migita \textit{et al}. (+0.10 \pm 0.01 mV K\textsuperscript{-1}).\textsuperscript{18} The measured value for Fc/15 mM Fc\textsuperscript{+} is slightly lower than that reported by Abraham \textit{et al}. (+0.154 mV K\textsuperscript{-1} in [Emim][NTf\textsubscript{2}])\textsuperscript{30} although those samples were ca. 30 fold more concentrated.

The \([\text{Fc}][\text{I}_3]\) system is more complex, and interestingly displayed a non-linear potential vs temperature trend. The \([\text{Fc}][\text{I}_3]\) also displayed consistently higher potential values that that of the individual ferrocene and iodine systems. The non-linear trend was present during heating and cooling, displaying no hysteresis, although it is important to note that it took ca. 1 h for these values to stabilise.

A maximum potential of ca. 24.4 mV was achieved by \([\text{Fc}][\text{I}_3]\) at \(T = 30\) K, which equates to an \(S_e\) of ca. 0.83 mV K\textsuperscript{-1}. Therefore the simple expedient of mixing Fc\textsuperscript{0} with I\textsubscript{3} results in an \(S_e\) value which (at \(T = 30\) K) equates to a 14-fold increase over \(\Gamma^-/\text{I}_3^-\) alone, and a 5-fold increase over the sum of the individual Fc/Fc\textsuperscript{+} and \(\Gamma^-/\text{I}_3^-\) redox couples, representing a significant synergy when the two are mixed to form \([\text{Fc}][\text{I}_3]\). Further increases in temperature resulted in some instability due to gradual evaporation of the volatile materials from the open U-tube, and at ca. 100 °C \([\text{Fc}][\text{I}_3]\) underwent irreversible decomposition to form iron iodide.

The same systems were investigated in a platinum-coated CR2032 battery casing, which provides 2 mm electrode separation (cf. 2 cm in the U-tube) and is more suited for power generation. The hermetically sealed casing also allowed higher temperature differences without any loss of the volatile redox compounds. In this case, the \([\text{Fc}][\text{I}_3]\) initial trend matched that of the U-tube but deviated at higher temperatures (\(S_e = 0.291 \pm 0.002\) mV K\textsuperscript{-1}, hollow symbols in Fig. 1(a)). Nevertheless, the \(S_e\) value still displayed the synergy discussed above. The \(S_e\) value of Fc/Fc\textsuperscript{+} was also measured in the battery casing and found to be 0.081 \pm 0.001 mV K\textsuperscript{-1} (cf. 0.100 \pm 0.002 mV K\textsuperscript{-1} in the U-tube), and all systems gave ca. 20% lower Seebeck coefficients in the battery casing. This is predominately due to a temperature drop from the heated exterior to the interior of the cell. In the U-tube (electrodes separated by 2 cm) it took ca. 1 h for the potential values to stabilise, whereas in the battery casing (2 mm separation) it took only a few minutes (Fig. S3, ESI). In the battery casing there is more significant thermal convection, whereas in the U-tube the temperature-induced change in the equilibrium in eqn (2) appears to become more significant over time. This mirrors observations in aqueous systems where varying the electrode separation and electrode orientation to gravity can either increase or decrease power according to the precise redox couple utilised.\textsuperscript{19} The power discharge characteristics of the systems were also investigated in the battery casing. As demonstrated in Fig. 2,
30 mM [Fc][I₃] generated significantly more power than 30 mM Fc/Fc⁺, largely due to the higher Sₑ value of the former. The precise charge carrier(s) involved in the [Fc][I₃] system is currently unknown. The power also refers to the steady-state conditions obtained after 10 min discharge. Given their differing diffusion coefficients12,13 and the possibility for both species to undergo Grothuss-like conductivity, further work is required to investigate what happens to these systems under extended discharge conditions.

While the temperature-dependant equilibrium in the ferrocium triiodide salt clearly has some influence, it does not explain the observed synergy resulting in the Sₑ increases. To investigate further, the ferrocene core was systematically altered by adding electron-donating (alkyl) or electron-withdrawing (alkanoyl) groups. Fig. 1(b) displays the potential differences measured for 45 mM I₂ and 30 mM of Fc (□), AcFc (▲), EtFc (●), DiBuFc (★) and DiBoylFc (♦) in [Emim][NTf₂]. Table 1 lists the Sₑ values of all redox couple species studied in this work, and it was found that 30 mM “[DiBoylFc][I₃]” possesses the highest Sₑ value of 1.67 ± 0.05 mV K⁻¹.

The trend in Sₑ values follows the order DiBoylFc > AcFc > Fc > DiBuFc > EtFc, which follows the anticipated electron density on the ferrocene centre from lowest to highest. The inversion of DiBuFc and EtFc is an exception to this trend, and is likely due to a subtle contribution from the alkyl chains to the overall entropy change.

In organic solvents, electron-donating groups on the ferrocene moiety (e.g. alkyl groups) shifts the equilibrium in eqn (2) towards the right, resulting in more triiodide product formed.16a,b This was investigated in [Emim][NTf₂] by cyclic voltammetry (CV), and the oxidation potentials of the substituted ferrocenes followed the expected trend in the IL (Fig. S5, ESI†). Comparison of the potentials relative to those of I⁺/[I₃]⁻ (Fig. S6, ESI†) and use of the Nernst equation allows us to estimate (see ESI† for full discussion) that a 2:3 mixture of DiBuFc and I₂ should exist as 99.9% [DiBuFc][I₃] (at 298 K); for DiBoylFc only ca. 10% exists as [DiBoylFc][I₃]. However, when ferrocene and iodine were mixed, the I⁺/[I₃]⁻ feature displayed a ca. 50 mV anodic shift, the [I₃]/I₂ process was unchanged, while the DiBoylFc/[DiBoylFc]⁺ peaks shifted ca. 50 mV cathodically. Additionally, currents for the mixed ferrocenium triiodide systems were higher than expected, demonstrating some ‘nonadditivity’ of faradaic currents; this has been previously noted when ferrocene is mixed with cobaltocenium in ionic liquids.20

Further demonstration of this non-linear behaviour was obtained by UV-Vis spectroscopy. Ferrocene-iodine charge transfer complexes are known in non-aqueous solvents,16d and extinction coefficient of I₂ and [I₃]⁻ are extremely structure- and solvation-sensitive.21 Fig. 4(a) displays the titration of I₂ with I⁻ (introduced as [Emim]I), yielding the characteristic absorption bands.16d,22 Stoichiometric quantities of I₂ and I⁻ resulted in quantitative formation of [I₃]⁻, as represented by the absorption maxima. Fig. 4(b) displays the same spectra of the UV-Vis absorption spectra of (a) 3 mM I₂ in presence of various [Emim]I concentrations, and (b) 3 mM I₂, 3 mM Fc/3 mM Fc⁺, 3 mM I⁻/3 mM [I₃]⁻ as well as 3 mM of the 5 [ferrocenium][I₃] samples. All in [Emim][NTf₂]. Pathlength 1 mm.
I₂ and I₂/I⁻, overlaid with the spectra for Fe₃[Fe[PF₆]₃] and all five “ferrocenium triiodides” salts (as 3 mM ferrocene and 4.5 mM I₂). All five ferrocene/iodine mixtures should have an absorbance below that of 3 mM [I₃]⁻, as they will introduce 3 mM [I₃]⁻ or less, based upon the equilibrium in eqn (2). As the absorbance of the ferrocenium triiodides exceed this in every case, the solvation environment of [I₃]⁻ in the IL is clearly altered upon introduction of ferrocene species. An interaction between ferrocene and iodine species will result in an additional entropic change upon a redox process occurring, relative to the individual species. These significant interactions can therefore explain the synergistic enhancement apparent in the ΔS coefficient. It is logical that the electron-deficient ferrocenes would display more significant charge-transfer complexation with the electron-rich and highly polarisable iodine species. This increased interaction mirrors the observed synergistic enhancement in the ΔS coefficient, with DiBoylFc predicted to interact with iodine the most and also displaying the highest ΔS coefficient of 1.67 mV K⁻¹. This ΔS coefficient compares favourably with the highest ΔS value ever reported in the IL [Emim][NTf₂], which was ca. 1.66 ± 0.02 mV K⁻¹ for 100 mM [Co⁹⁺[bpy]₃][NTf₂]₂⁻. The reported [Co⁹⁺[bpy]₃][NTf₂]₂⁻ was deliberately chosen to result in the most significant entropy change upon changing oxidation state, with purely the cationic component contributing. In the case of the ferrocenium triiodides, we have demonstrated that two common, widely available, relatively poor thermoelectrochemical redox mediators (by conventional knowledge and approaches) can be combined to result in an excellent system when combined.

In conclusion, this work has investigated the thermoelectrochemistry of mixed redox couples (ferrocene and iodine) for the first time. These two couples were found to interact even in dilute solution, resulting in a synergistic thermoelectrochemical response. This likely stems from observed charge-transfer complex formation, which must introduce additional entropic changes when one component undergoes a redox process, additional to simple restructuring of the surrounding ionic liquid. This work therefore represents a new approach towards the development of thermoelectrochemical systems, with the value of 1.67 mV K⁻¹ for [DiBoylFc][I₃] being one of the highest reported ΔS value for a solute in the ionic liquid [Emim][NTf₂].

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Notes and references